1. Enzymes belong to which class of compound: 
   Proteins

2. The helical structure of protein is stabilized by 
   Hydrogen

3. Enzymes have 
   Polypeptide

4. Function of enzymes in biological system is 
   Catalyze biochemical action

5. Night blindness is produce due to deficiency of ... 
   VITAMIN A

6. Zwitter ion exist in which of the following 
   a) Alanine 
   b) Glycine HCL 
   c) Both a and b 
   d) None

7. IR determine ----------------- bond 
   kind of bond nature of bond

8. Marsh gas is 
   ...CH4

9. Solvent used in IR Spectroscopy is:
a) **CCL4**  
B) **CH30H**  
C) **WATER**  
D) **CH3CH2OH**

10. **1700------1760 PEAK In IR Spectroscopy is:**  
   Carbonyl group

11. **NMR use ----------------source of radiation**
   
   a) **UV**  
   B) **Radio waves**  
   c) **micro waves**  
   d) **visible rays**

12. **Existence of nucleus is confirmed by............**  
   Millikian method or  
   **Alpha scattering by thin metal foil**

13. **What is false about electron**  
   (Not conform but it should be d)  
   a) Electron is a particle  
   b) Wave is associated with electron  
   c) When electron jump from orbital its energy c..  
   **d) oriented by magnetic field**
14. Which one is occur in whole number always
   a) **Atomic number**
   b) Molecular weight
   c) Equivalent weight

15. A metal has an electronic configuration is M+2 has 2,8,14 and its atomic weight is 56. the number of neutron is:
   30

16. Least soluble in Water Is
   a. Ethano
   b. Phenol
   c. Carboxylic acid
   d. **benzene**

17. The EAN Value of Ni(CO)4 is:
   a) 36
   b) 54

18. The ration of C:Si in their neutron ration with atomic weight having 12 and 28 respectively
   (3:7)

19. Phenolphthalein is colourless in acidic medium due to:
20. The weakest base is
   a) H-
   b) Cl-
   c) HCO3-
   d) OH-

21. Which compound decolorize KMnO4 solution but not give test with AgNO3 solution
   a) Ethylene
   b) Acetylene
   c) Ethane

22. Which one of the following does not give Lassaignes test for nitrogen
   A. Phenylhydrazine
   B. Glycine
   C. Urea
   D. Azobenzene

23. Which one the following compound does not explode:
   a) trinitro toluene
   b) b)2-amine toluene

24. Strongest oxidizing agent is supported by---------
   a. Low dissociation energy
   b. high electron affinity
   c. low ionization energy
25. the order of oxidizing agent is
   I<Br<Cl<F

26. inversion of cane sugar into glucose and fructose is ?(not conform)
   a) slow process
      fast process
   b) spontaneous process
      instantaneous process

27. Rate of chemical reaction is---------when it proceed
   a. decrease
   b. increase

28. when reactants are added then rate of reaction is
   decrease
   increase (right answer)

29. which one of the following metal has highest electrical conductivity
   A. gold
   B. silver
   C. aluminium
   D. copper

30. preparation of ice cream is

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a) colloids

b) excessive cooling
c) coagulation
d) peptization

31. colloids are example of ---------------equilibria
   a. homogenous equilibria
   A. heterogenous equilibria
   b. suspension
   c. none of these

32. pure water is nonconductor of electricity
   A. water is neutral
   B. water is almost unionized

33. which one of the following aqueous solution conductor of electricity
   C. Glucose
   D. HCl

34. Which is correct statement
   A. Catalyst increase the rate of forward reaction
   B. Catalyst increase the rate of backward reaction
   C. Catalyst influence on the rate of forward reaction and backward reaction equally
   D. Catalyst decrease the rate of forward reaction

35. Pick a right statement:
   Order of reaction is in fraction
36. The unit for formation of ammonia from \( \text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3 \) is
\[ \text{Mol} - \text{litre} \]

37. The unit of active mass is
A. Gram atom per litre
B. Gram molecule per litre
C. Gram litre

38. Colloidal particles are
a) negative
b) positive
c) neutral
d) may be positive or negative

39. Who classify element first
a) lother meyer
b) doberiener
c) newlands
d) mendeleevs

40. Molecularity of reaction is always in
\[ \text{whole number} \]

41. The chemical which destroys ozone layer is
\[ \text{CFC}_3 \]

42. Physisorption occurs at
temperature low

43. if dispersion medium is gas its example is

Aerosol

44. pressure vs temperature equation is called

A. Langmuir equation
B. Fredluck equation
C. Bet equation
D. None of these

45. Resonationg structure caanot be

separated

46. Nickel is use in the hydrogenation of

vegetable ghee

47. Colloids shows

Brownian movement

48. Methyl group is Ortho and p-directiong bcz........

a) E donating effect of methyl group inductively
b) E donating effect of methyl group resonatingly
c) E donating effect of methyl group hyperconjugation (right answer)
d) All of above

49. Chloro group is ortho and para but

electron withdrawing
50. To avoid maximum yield for a reversible reaction to occur
   A. Glass vessel
   **B. Closed vessel**
   C. Open vessel

51. In benzene, there are -------------- type of carbon atom
    **A. One type**
    B. Two type
    C. Three type
    D. Four type

52. In benzene, the number of pi electron is
    A. 3 e
    B. 4e
    **C. 6e**
    D. 4e

53. In benzene, all carbon atoms are------------hybridization
    A. Sp3
    **B. Sp2**
    C. Sp
    D. None

54. According to Lechatelier principle in a reversible reaction between solid and liquid, the amount of 
    heat added
    A. Decrease the concentration of solid
    B. increase the concentration of solid
C. Decrease the concentration of liquid
D. Decrease the concentration of liquid

55. The fastest reaction occur in which of the following:
A. Rusting of iron
B. Burning of coal
C. The AgCl formation by the reaction of AgNo3 and Nacl

56. The amount of energy required to separate nucleons is called
A. Ionization energy
B. Binding energy

57. The enthalphy of heat of formation of sodium ion from its atom is known as
A. Enthalphy of atomization
B. First ionization energy
C. Enthalphy of combustion

58. The brown vapour is formed of which element when MnO2 and conc. H2SO4 is added in them
A. NO2
B. Cl2
C. Br2
D. I2

59. Stability of nucleus is due to
A. Proton and neutron
B. Proton and electron
C. Neutron and electron
60. Difference between particle mass and real mass is called ...........

**Mass defect**

61. A gas has a volume of 2 litre at S.T.P at constant pressure, the new volume is 4 litre at which temperature:
   
   A. 50c
   B. 100c

62. The difference between the crystalloids and colloids

   A. **Particle size**
   
   B. Colloidal size
   
   C. Diffusion
   
   D. True solution

63. The gas which shows deviation from ideal behavior is:

   A. Helium
   B. Hydrogen
   
   C. **Ammonia**
   
   D. Trichloro methane

64. Some substance shows scattering of light but they can pass through ordinary filter paper is ............

   A. **Colloids**
B. Suspension
C. True solution
D. None

65. In the lime kiln, CaCO3..................>CaO+CO2 proceed to complete because...........
   A. CO2 ESCAPPE
   B. CaO remove
   C. Low Temperature
   D. High Pressure

66. Which one of the following equation is not affected by change in pressure
   A. \(\frac{1}{2}N_2 + \frac{1}{2}O_2 \rightarrow NO\)
   B. \(P_2O_5 \rightarrow P_2O+O_2\)
   C. \(N_2+3H_2 \rightarrow 2NH_3\)

67. A Reaction approaches go farthest completions
   A. \(K=10^2\)
   B. \(K=10^1\)
   C. \(K=10\)
   D. \(K=10^{-2}\)

68. Isothermal thermal system is one in which
   **\(\Delta T=0\)**
   \(\Delta E=W\)
   \(V=0\)
   \(E=0\)
69. Temperature of the system decreases in 
**adiabatic expansion**
adiabatic compression
70. System in which no change in temperature is 
Isochoric
isobaric
**isothermic**
71. Cyclic structure of benzene is given by
**Kekule**
Doberiener
72. Which is least soluble in water
A. iron
**B. Benzene**
C. Benzoic acid
D. Phenol
73. Heterogenous catalysis is based on
**a. Adsorption**
A. Absorption
B. Sorption
C. Desorption
74. Entropy of the universe is
**Always increasing**
Decreasing
Constant
75. Which indicator is not used in conical flask
A. Self indicator
B. External indicator  (right answer)
C. Internal indicator
D. Mixed indicator

76. Toluene is prepared from benzene
A. Fridel craft reaction
B. Perkin reaction
C. Wurtz reaction

77. Catalyst used in fridel craft alkylation is
A. AlCl₃
B. FCl₃

78. Order of reaction is.........
Can be determined from equation
Can be determined from experiment

79. Oxidation number of P in KH₂PO₂ is
+1
+3
+5
+2

80. Molten KCl conduct electricity because of
Free ions
Free electron

81. which individual units loose identity
   a. Complex salts
   b. Simple salts
   c. Compound salts
   d. Double Salt

82. When a multidentate ligand is surrounded it is known as
   Coordination sphere
   Coordination complex

83. Group no. represents
   Same electronic configuration
   same number of electron in valance shell Chelates

84. Entropy decreases in melting of ice rusting of iron

85. In this reaction $6\text{NaOH} + 3\text{Cl}_2 \rightarrow 5\text{NaCl} + \text{NaClO}_3 + 3\text{H}_2\text{O}$
   A. Chlorine is reduced
   B. Chlorine is oxidized
   C. Disproportionation reaction

86. Aqueous solution is electrical conductor MgCl$_2$
87. According to lewis concept,
   bases are electron pair donor
88. According to lowry-bronsted concept
bases are proton acceptor

89. Which one of the following aqueous solution of oxide and chloride give same test with litmus solution.
    A. Phosphorous
    B. Sodium
    C. Magnesium

90. The element of group 1B is known as coinage metal.

91. Oxidation of which gives positive DNP but not Fehling's test
    A. CH3CH2CH2CH2OH
    B. CH3CH(OH)CH2CH3
    C. CH3CH(CH3)CH2CH2OH
    D. (CH3)3COH

92. Product of hydrolysis of CH3CO2C3H7 with NaOH gives
    A. CH3COONa +C3H7OH

93. Fluorescene, rhodamine, are example of---------indicators.
    A. Acid base indicator
    B. Adsorption indicator
    C. Absorption indicator